Symbols in coordination equilibria.

Elements

ML_n	symbolises a coordination compound in equilibrium in aqueous solution To simplify charges has been omitted.
n	is an integer ($0 \le n \le N$) : it means the number of (monodentate) ligands - L – which are bound to the metal ion M.
N	The maximum coordination number is N.
ML_n	is therefore a short notation for $ML_n(H_2O)_{N-n}$ (charges omitted). M may be called metal ion, central ion or central atom

Concentration

$[ML_n]$	molar concentration of the species ML_n
C_x	stoichiometric concentration of X - the molar concentration of X in total bound to the metal ions or free. Is often called the total concentration of X. For typographical reasons it may also be written as C(X).
C_M	Stoichiometric concentration of metal ion
C_L	Stoichiometric concentration ofligand

The degree of formation

α_n	The degree of formation of ML_n (For typographical reasons it may also be written as α(n)). $\alpha_n = \frac{[ML_n]}{C_M}$
$\bar{n}$	The degree of formation of the system - (is pronounced n-mean). $\bar{n} = \frac{\sum n \cdot [ML_n]}{C_M} = \sum n \cdot \alpha_n$

Stability constant

K_n	the step wise stability constant for the step $ML_n + L \leftrightarrow ML_n$ $K_n = \frac{[ML_n]}{[ML_{n-1}] \cdot [L]} = \frac{\alpha_n}{\alpha_{n-1} \cdot [L]}$
β_n	$\beta_n = \frac{[ML_n]}{[M] \cdot [L]^n} = \frac{\alpha_n}{\alpha_0 \cdot [L]^n}$ overall stability constant $\beta_n = \prod_0^n K_n$